

1. Name the following :. (5)
- The type of bonding in Ammonia molecule.
 - The ion which combines with water molecule to form hydronium ion.
 - Process of formation of ions from molecules which are not in ionic state.
 - The bond formed between dissimilar atoms having different electronegativity.
 - The energy released when an electron is added to a neutral gaseous isolated atom to form a negatively charged ion.
2. Give reason :. (3)
- Ionisation Potential increases across a period ,from left to right.
 - Ionic compounds have high melting points.
 - Hydrogen chloride can be termed as a polar covalent compound.
3. Use electron dot diagram to show the formation of the stable positive ion formed when acids dissolve in water. (2)
4. Out of the two elements X and Y which has bigger size ? Why ?(2)
- X has atomic number 12 and atomic mass 24
 - Y has atomic number 11 and atomic mass 23.
5. The following table represents three elements and their atomic numbers. With reference to this, answer the following using only alphabets given in the table. (3)

Element	Atomic Number
P	20
Q	7
R	10

- Which element combines with hydrogen to form a basic gas ?
- Which element has an Electron Affinity zero ?
- Name the element which forms an ionic compound with chlorine.

6. Draw the electron dot diagram for the compounds given below :. **(2)**

a) Calcium Oxide

b) Water molecule

7. The elements of one short period of the Periodic Table are given below in order from left to right.

Li Be B C O F Ne

i) To which period do these elements belong?

ii) One element of this period is missing. Which is the missing element and where should it be placed?

iii) Place the three elements Fluorine, Beryllium and Carbon in the order of increasing electronegativity.